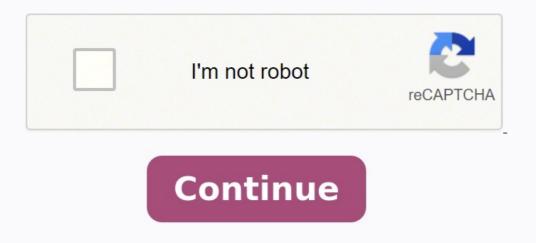
Thermodynamics and statistical mechanics shell solution manual



Responsible M. Scott Shell (University of California, Santa Barbara). Cambridge Publications: Cambridge University Press, 2015. Copyright information © 2015 Physical description xxii, 476 pages: illustrations; 26cm. Series Cambridge Series in Chemical Engineering. Start with a Call: Thermodynamics and Statistical Mechanics: An Integrated Approach M. Scott ShellCambridge University Press, 2015 CHE 210A: Thermodynamics and Statistical Mechanics (F15) UCSB Department of Chemical Engineering II 3301 Website: www.engr.ucsb.edu/~shell/che210a/ Course Description Thermodynamics and statistical mechanics are, at their most fundamental level, the sciences of equilibrium. Thermodynamics takes a macroscopic, molecular level. This course brings together the concepts and features and properties of general equilibrium, relationships and constraints between thermodynamic variables, and solids. Particular attention is paid to understanding the relationship between macroscopic volume behavior and molecular interactions. Instructor Reader Professor M. Scott Shell Office: Engineering II 3321 Phone: (805) 893-4346 Email: M. Scott Shell, Cambridge University Press, 2015 Other Recommended Texts Molecular Driving Forces: Statistical Thermodynamics in Biology, Chemistry, Physics and Nanoscience, Ken Dill and Sarina Bromberg, Garland Science, 2010 An excellent modern introductory text on the topic, especially good - suitable for self-study. Statistical Mechanics, Donald A. McQuarrie, University Science Books, 2000 (2nd ed.) A classic text in statistical mechanics with a clear and concise presentation. Credit M. Scott Shell (University Press, 2015. Copyright information © 2015 Physical description xxii, 476 pages: illustrations; 26 cm. Cambridge Chemical Engineering Series. Start with a Call: Thermodynamics and Statistical Mechanics: An Integrated Approach M. Scott ShellCambridge University Press, 2015 CHE 210A: Thermodynamics and Statistical Mechanics (F15) UCSB Department of Chemical Engineering II 3301 Website: www.engr. ucsb.edu/~shell/che210a/ Course Description At a basic level, thermodynamics and statistical mechanics are the sciences of equilibrium. Thermodynamics takes a macroscopic volumetric view, while statistical mechanics defines the principles of equilibrium at the microscopic, molecular level. This course will bring together both perspectives to discuss general features and properties of equilibrium, relationships and limitations between thermodynamic variables, and simple models of gases, liquids, and solids. Particular attention will be paid to understanding the relationship between thermodynamic variables, and simple models of gases, liquids, and solids. Scott Shell Office: Engineering II 3321 Phone: (805) 893-4346 Email: M. Scott Shell, Cambridge University Press, 2015. Other Recommended Reading Molecular Driving Forces: Statistical Thermodynamics in Biology, Chemistry, Physics, and Nanoscience, Ken Dill and Sarina Bromberg, Garland Science, 2010. An excellent modern introductory text on the topic is especially good - suitable for self-study. Statistical Mechanics, Donald A. McQuarrie, University Science Books, 2000 (2nd ed.) A classic text in statistical mechanics with a clear and concise presentation.subject material. It serves as both an introduction to the subject and a reference work. This book is also often used in CHE 210B. An Introduction to Statistical Thermodynamics, Terrell Hill, Dover Books, 1987. This inexpensive paperback presents the foundations and early theoretical advances in statistical mechanics. Hill discusses many subtleties missed by other texts and provides detailed conclusions. This is an essential reference for anyone involved in research broadly related to molecular science. Course Policy 1. Final grades are calculated based on: 20% classroom interaction, 30% midterm exam, 50% final exam. 2. Students must adhere to the highest ethical standards in all aspects of the course. This involves working on problems independently rather than looking online or elsewhere for external solutions or answers. How to succeed in this course Thermodynamics and statistical mechanics can be difficult subjects to understand. A set of healthy habits will make them clear and consistent: 1. You, not me, are responsible for learning. At graduate level, you must take full responsibility for a thorough understanding of the subject, independent of specific homework and lectures. Read the course materials carefully, cover the topics in detail and, if necessary, look for other resources to help you understand. It's about knowing thermodynamics and statements you do not understand. Write them down, take notes, keep track of what is causing you difficulty and confusion. Then look for answers. Complete the tasks a second time. Please see another tutorial. Ask a colleague. Time to visit the office. If you don't understand it well enough to teach it to someone else, you haven't achieved your goal. 3. Be an independent learner. Work independently on problems Try to remove obstacles by using different approaches, working on different examples or reading different texts. Then contact colleagues to discuss the best approach. Don't take issue with others at the outset, or you'll miss things you don't understand. 4. Ask yourself, "What would happen to this problem if X changed to Y?" Challenge yourself to think of possible variations that are not mentioned in the lectures and assignments. 5. Invest the time. Thermodynamics requires quality practice to become familiar with concepts, calculations, and mathematical manipulations. Don't be afraid to include more examples in the book or find new books to consult. Your goal should not be to master the assignments, but to master the topic. 6. You will fight. This is a challenging course and you will only understand the material if you engage in it by reading, thinking and working. Don't expect to know the answer right away and don't compare yourself to other students, as you will find that everyone has a different experience. Your job is to learn and understand, not to turn the corner. Homework problems assigned 2.6 3.6 3.11 3.16 4.9 5.13 6.10 6.25 7.24 8.7 8.8 (g) 8.21 9.21 10.11 11,6,10 12.7 12.11 12.28 12.1 14,13 15.8 (C, G, H, H 15,16 16,8,11 16,23 19.12 18,20 19.20 Code. UC Policy 102.23 expressly prohibits students (and anyone else) from recording lectures or discussions and distributing or selling lecture notes and any other course materials without the prior written permission of the instructor (see /2710530 /PACAOS -100). Students may take notes for their own private educational purposes only. 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W How the macroscopic world works 3.6, 3.9, 3.11 10/2 F no class â Graduate Symposium 4 10/5 M fundamental equations 3.16, 4.3, 4.10 5 10/7 W first law and reversibility 5.9, 5.13, 6.10 6 F potential transformations and other legends 6.20, 6.25 7 10/12 M Maxwell's relations and measurable quantities 7.6, 7.16, 7.24 8 10/14 W gases 8.7, 8.8g, 8.21 9 10/ 16 F109,1,13 equilibria, 101,101 10,12 11 10/21 W solutions 12,11, 12,21 13 10/26 M solids 12 .28, 12.29 14 M 10/21/M 3.1 11/2 M canonical division function 15.8cf, 15.16 17 11/4 W canonical division function 15.8cf, 15.16 17 11/4 W canonical division function 16.1, 16.8, 16.11 18 11/6 F fluctuations 16,12, 16,23 W 11/9 M no class/SS19 trip no class 13 F Statistical mechanics of classical systems 18.11, 18.14, 18.20 21 11/18 W Other routes M 1/20 choirs M3119 class. 19.16 23 11/25 W reaction coordinate u19.20, 19.21 27.11. F without lessons â holiday 30.11. M Methods of molecular modeling 2.12. 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